

ĐÁP ÁN KIỂM TRA HỌC KỲ II - NĂM HỌC 2020 - 2021

Môn : HÓA HỌC 12

PHẦN I. TRẮC NGHIỆM (24 câu - 6 điểm)

Đáp án mã đề: 401

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PHẦN II. TỰ LUẬN (4 câu - 4 điểm)

Câu	Đề	Đáp án chi tiết	Điểm thành phần
1 (1 điểm)	Kim loại sắt tác dụng được với dung dịch muối nào sau đây? Fe(NO ₃) ₃ , ZnCl ₂ , AgNO ₃ , CuSO ₄ . Viết phương trình minh họa?	$\text{Fe} + 2\text{Fe}(\text{NO}_3)_3 \rightarrow 3\text{Fe}(\text{NO}_3)_2$ $\text{Fe} + \text{ZnCl}_2 \rightarrow \text{không xảy ra}$ $\left\{ \begin{array}{l} \text{Fe} + 2\text{AgNO}_3 \rightarrow \text{Fe}(\text{NO}_3)_2 + 2\text{Ag} \\ \text{Hay } \text{Fe} + 3\text{AgNO}_3 \rightarrow \text{Fe}(\text{NO}_3)_3 + 3\text{Ag} \end{array} \right.$ $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$	0,25x4
2 (1 điểm)	Cho hỗn hợp CuO, Fe ₂ O ₃ , Al ₂ O ₃ tác dụng với CO dư, nhiệt độ cao. Viết phương trình phản ứng (nếu có).	$\text{CuO} + \text{CO}, t^0 \rightarrow \text{Cu} + \text{CO}_2$ $\text{Fe}_2\text{O}_3 + 3\text{CO}, t^0 \rightarrow 2\text{Fe} + 3\text{CO}_2$ <p>- Nếu viết 3 phương trình -0,25đ - Thiếu điều kiện hay cân bằng -0,25đ</p>	0,5x2
3 (1 điểm)	Cho từ từ bột Fe vào 50ml dung dịch CuSO ₄ 0,3M, khuấy nhẹ cho tới khi dung dịch mất màu xanh. Tính <u>khối lượng</u> bột Fe đã tham gia phản ứng?	Mol CuSO ₄ = 0,015 mol $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu} \dots\dots\dots$ 0,015 0,015 $m_{\text{Fe}} = 0,015 \times 56 = 0,84\text{g} \dots\dots\dots$	0,25 0,25 0,25 0,25
4 (1 điểm)	Cho 5,6 gam Fe tác dụng với 250ml dung dịch AgNO ₃ 1M lắc kỹ để phản ứng xảy ra hoàn toàn thu được dd A và m (g) chất rắn. Tính giá trị m?	$n_{\text{Fe}} = 0,1 \text{ mol}$ $n_{\text{AgNO}_3} = 0,25 \text{ mol}$ lập tỉ lệ mol AgNO ₃ / Fe = 2,5 → 2 muối $\left. \begin{array}{l} \text{Fe} + 2\text{AgNO}_3 \rightarrow \text{Fe}(\text{NO}_3)_2 + 2\text{Ag} \\ a \qquad 2a \qquad \qquad \qquad 2a \\ \text{Fe} + 3\text{AgNO}_3 \rightarrow \text{Fe}(\text{NO}_3)_3 + 3\text{Ag} \\ b \qquad 3b \qquad \qquad \qquad 3b \end{array} \right\}$ $n_{\text{Ag}} = n_{\text{AgNO}_3} = 0,25$ $\rightarrow m_{\text{Ag}} = 0,25 \times 108 = 27\text{g} \dots\dots\dots$	0,25 0,25 0,25 0,25